

## Worksheet 7: Molar Mass Practice

Find the Molar Mass of the following compounds:

- |   |  |
|---|--|
| a. $\text{H}_2\text{O}_2$                                 | h. $\text{Al}(\text{CN})_3 \cdot 5\text{H}_2\text{O}$  |
| b. $\text{CaCO}_3$  | i. $\text{CoC}_2\text{O}_4 \cdot 10\text{H}_2\text{O}$ |
| c. $\text{Sb}_3(\text{PO}_3)_5 \cdot 3\text{H}_2\text{O}$ | j. $\text{Sc}_2(\text{SO}_3)_3$                        |
| d. $\text{BaC}_2\text{O}_4$                               | k. $\text{NiC}_2\text{O}_4 \cdot 3\text{H}_2\text{O}$  |
| e. $\text{Nb}(\text{CN})_5 \cdot 2\text{H}_2\text{O}$     | l. $\text{Pb}(\text{Cr}_2\text{O}_7)_2$                |
| f. $\text{Te}_3\text{I}_8$                                | m. $\text{Co}_3(\text{BO}_3)_2$                        |
| g. $\text{Pb}(\text{SO}_3)_2$                             | n. $\text{AlN} \cdot 7\text{H}_2\text{O}$              |

Write the formulas for the following compounds. Find the molar mass of the compounds.

- Silicon tribromide
- Tin (II) phosphide
- Aluminum permanganate monohydrate
- Cobalt (III) sulphate nonahydrate
- Tetraarsenic diiodide
- Lead (IV) borate
- Phosphorus hexabromide
- Sulphurous Acid
- Ammonium permanganate nonahydrate
- Tetrasilicon hexaoxide